

PERIODIC TABLE
CBSE/WBCHSE/JSC BOARD
ASSIGNMENT SET 1

1. State Modern Periodic Law. Name the two elements of first period.
2. Two element X, Y and Z belong to 17th group but to 2nd, 3rd and 4th period respectively. Number of valence electrons in Y is 7. Find the number of valence electrons in X and Z.
3. How does the metallic character of the elements vary (i) in a group (ii) in a period of the modern periodic table?
4. Na, Mg, Al and P belong to 3rd period but are placed in first, second, thirteenth and fifteenth group. Number of shells occupied in Mg is three. What is the number of occupied shells in Na, Al and P. Give reason for your answer.
5. What is the number of elements in first, second and third period of the periodic table? Give reason for your answer.
6. "Elements in Periodic Table show periodicity of properties" List any 4 such properties.
7. (a) Atomic number of Mg and Al are 12 and 13 respectively. Write their electronic configuration. (b) Mention the period of the Modern Periodic Table to which the above two elements belong. Give reason for your answer.
8. (a) State the Mendeleev periodic law. (b) What is the total number of periods and groups in Modern Periodic Table.
9. Nitrogen (atomic number 7) and phosphorous (atomic number 15) belong to group 15 of the Periodic Table write the electronic configuration of these two elements. Which of these will be more electronegative and why
10. (a) Element Y, with atomic number 3 combines with element A, with atomic number 17, what would be the formula of the compound? (b) What is electronic configuration of element with atomic number 29? (c) What will be its valency?